

REACTION KINETICS CATALYST AND
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DESCRIPTION: The purpose of this study is to investigate the effect of catalyst concentration on the reaction rate of the reaction between hydrogen peroxide and potassium dichromate in the presence of various metal ions. The reaction is a redox reaction in which hydrogen peroxide is oxidized to oxygen gas and water, while the metal ions are reduced to their corresponding lower oxidation states. The reaction is carried out in a glass beaker and the liquid phase is stirred continuously with a magnetic stirrer. The reaction mixture is a clear, colorless solution. The reaction is exothermic and the temperature of the reaction mixture increases during the course of the reaction. The reaction is first order with respect to the concentration of hydrogen peroxide and zero order with respect to the concentration of the metal ions. The reaction rate is determined by measuring the volume of oxygen gas evolved over a period of time. The reaction rate increases with increasing catalyst concentration and the reaction is complete within a few minutes.

RESULTS: The reaction rate increases with increasing catalyst concentration. The reaction is first order with respect to the concentration of hydrogen peroxide and zero order with respect to the concentration of the metal ions. The reaction rate is determined by measuring the volume of oxygen gas evolved over a period of time. The reaction rate increases with increasing catalyst concentration and the reaction is complete within a few minutes.

CONCLUSION: The reaction rate increases with increasing catalyst concentration. The reaction is first order with respect to the concentration of hydrogen peroxide and zero order with respect to the concentration of the metal ions. The reaction rate is determined by measuring the volume of oxygen gas evolved over a period of time. The reaction rate increases with increasing catalyst concentration and the reaction is complete within a few minutes.

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