style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility ?B? style.visibility =?`B?   ?B?
style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility ?B? style.visibility =?`B?   ?B?
style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility ?B? style.visibility =?`B?   ?B?
style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility ?B? style.visibility ?B? ppt_x ?B?
ppt_y ?B? style.visibility P T'Y \ 0

style.visibility =?`B?  ?B? style.visibility ?B? style.visibility ?B? style.visibility ?B? style.visibility ?B? style.visibility =?`B? ?B? style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility =?`B? ?B? style.visibility ?B? ppt_x ?B? ppt_y ?B? style.visibility ?

$H_2(g)+1/2O_2=H_2O(L)$?rHm?(298K) == $-$ 241.8 kJmol^{-1}

$H_2(g)+1/2O_2(g)$

$H_2O(1)$

$2H+O(g)$

$2H+O(g)$

$?Hm?=?H_1?+?H_2?+?H_3?$

298K

$C(s) +O_2(g) \rightarrow CO_2(g)$; ?r1Hm? $-$ 393.51 kJmol^{-1}

$CO(g) + 1/2O_2(g) \rightarrow CO_2(g)$; ?r2Hm? $-$ 282.99 kJmol^{-1}

$C(s) + 1/2O_2(g) \rightarrow CO(g)$

?r3Hm? ?r1Hm? ?r2Hm?+?r3Hm? ?r3Hm? ?r1Hm?$-$?r2Hm? $-$110.52 kJmol^{-1}

298K

$Fe_2O_3(s)+3CO(g)=2Fe(s)+3CO_2(g)$?r1Hm? $-$ 25 kJmol^{-1}

$3Fe_2O_3(s)+CO(g)=2Fe_3O_4(s)+3CO_2(g)$?r1Hm? $-$ 47 kJmol^{-1}

$Fe_3O_4(s) +CO(g) =3FeO(s)+CO_2(g)$?r1Hm? 19 kJmol^{-1}

$FeO(s) +CO(g)= Fe(s)+CO_2(g)$

1 X 3 2 $-$2 X 3 ?Hm?=3 X $-$25 $-$ $-$ 47 $-$2 X 19? 0

style.visibility =?`B? ?B? style.visibility ?B? style.visibility =?`B? ?B?
style.visibility ?B? style.visibility =?`B? ?B? style.visibility ?B? style.visibility ?B?
style.visibility =?`B? ?B? style.visibility ?@ABCDEFGHIJKLMNOPQRSTUVWXYZ[\]
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